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<p>(54) Title: SULFIDE SCAVENGING PACKAGING MATERIALS (57) Abstract <p>A packaging material, which is capable of absorbing trace amounts of volatile odorous sulfur compounds from food, is made from a structural support material which is preferably a polymer film, and a defined molecular sieve. The packaging material may be in the form of a <u>film</u> which forms at least <u>one layer</u> of a <u>film package</u>, or it may be in the form of a cap liner, or it may be in the form of an <u>insert</u>. The molecular sieve may be defined either by its <u>measured ability to adsorb hydrogen sulfide</u>, or by certain physical and chemical structural parameters of the sieve. These are pore-opening size, Si/Al atomic ratio, pore-structure dimensionality, and basicity.</p></p>		

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TITLE

SULFIDE SCAVENGING PACKAGING MATERIALS

BACKGROUND OF THE INVENTION

Field of the Invention

5 The present invention relates generally to packaging materials which enhance food freshness by removing hydrogen sulfide and other odorous sulfur products from certain foods. More specifically, the packaging materials of the present invention comprise a defined molecular sieve enclosed by a polymer film or sheet or other material, or embedded within a polymer film or
10 sheet or other material. The packaging materials may be used as a packaging structure (i.e., a package itself such as a film package), a layer of a film package, or a package insert such as a coupon or liner.

Description of Related Art

15 In food packaging, the principle emphasis in relation to odor removal has been with the packaging material itself. Many polymeric packaging materials produce odors which can impart an undesirable odor to the food. Various additives have therefore been added to the polymer to remove its odor. However food itself, even while it remains in good condition, can produce undesirable odors, and some of the odor producing substances have
20 such a low odor threshold that trace amounts are undesirable. In addition, trace amounts of such substances can produce an off-taste in food.

 The proteins and free amino acids of meats, particularly poultry, break down on heating to form a variety of volatile compounds. Among these are sulfur-containing compounds, including hydrogen sulfide, mercaptans,
25 sulfides and disulfides. These sulfur compounds can discolor cans containing such food, and there has been an abundance of work directed to can coatings to prevent this discoloration. However, a more important aspect is the presence of the odor, and often an associated taste, in the food itself. The odor threshold of some of these is very low (10 micrograms/liter or less), so the presence of even
30 small amounts can, as noted, be very objectionable well before any real deterioration of the food has occurred.

Many materials can remove sulfides by reaction with them. Metal oxides, particularly zinc oxide have been used. Industrially, removal of sulfides in various processes is often carried out by reaction. In general, using
35 chemical reactions to remove sulfides is fine in chemical processes, but not very

suitable for food use, both because of toxicity concerns and the difficulty of controlling any such process. Adsorption is an alternative to reaction. Many adsorbers are unsuitable for contact with food. However, while several adsorbents may be entirely suitable for contact with food, there appears to be no work directed specifically to removing trace amounts of sulfur-compound related odors from food with their use.

Packaging materials capable of scavenging specific decomposition products from food have been disclosed in U.S. Patent No.: 5.284.892 (Brodie and Visioli). Specifically, aldehydes are scavenged by a polyalkylene imine of greater than 800 molecular weight. The imine is incorporated into a packaging film or sheet. Similar scavengers are disclosed in allowed Patent Application, serial number 08/176.740, which extends their use to packaging structures which include liners, inserts and the like. Polyalkylene imines are unsuitable for direct food contact, and so when used, the required packaging structure has to ensure that the imine does not contact the food.

Molecular sieves are materials of varying composition which have a three dimensional cage like structure which form channels with defined mouth opening (aperture or pore-opening) sizes. They may be naturally occurring materials, modified natural materials, or essentially synthetic. They include silicas, metalloaluminates, aluminophosphates and various others. The silicas may be subdivided into titanosilicates, aluminosilicates, gallosilicates, ferrisilicates, borosilicates, chromosilicates and others. Among the aluminosilicates are the zeolites. Specific zeolites have been used for different purposes including catalysis, sieving and adsorption. Typically, zeolites are defined by parameters such as Si/Al ratio, their pore-opening size and structure, acidity, the cations present and others.

Clays may also be aluminosilicates also containing magnesium, but typically have a laminar structure, not a three dimensional structure.

Japanese published application, J61.120.638A discloses an adsorbent composite for packaging consisting of a polyolefin related resin and an adsorbent. The adsorbent is broadly disclosed to be any of a diverse range of materials including silica gel, activated alumina, acid clay, activated clay, zeolites or active carbon. Uses include dehumidifying and deodorizing pharmaceuticals, foods, or precision machines.

U.S. Patent No. 4,795,482 (Gioffre et al.) discloses a process for removing odors from an 'environment', using certain specific hydrophobic, high silica zeolites and so-called silica polymorphs. Odors removable include those caused by a vast range of volatile materials including certain acids, aldehydes, organic nitrogen contain compounds and sulfur compounds including mercaptans and sulfides. Applications include powders, sprays, pads, creams, mouthwash, and non-woven tissues for treatment of odors in bathrooms, kitchens, refrigerators etc.

U.S. Patent No. 5,011,019 (Sato et al.) discloses a packaging structure for packaging medicines which may emit foul odors caused by substances such as carbonic acid, trimethylacetic acid and hydrogen sulfide. The package enclosing the medicine has a metal press-through side, and a multilayer polymer side, one layer of which is an polyolefin layer which incorporates deodorizing materials. The deodorizing materials mentioned as suitable include inorganic metal salts, favonoid, and unspecified molecular sieves.

There remains a need in the art for a packaging material, which can be used safely in contact with food, used either as a package itself, or as a packaging insert, specifically designed to remove hydrogen sulfide, mercaptans, sulfides and other odorous sulfur compounds from the specific foods which generate this odor, thus increasing shelf life.

SUMMARY OF THE INVENTION

The present invention is directed to novel packaging materials which form packaging structures or package inserts capable of adsorbing sulfur-compound odors generated by certain foods. The packaging material comprises a particulate aluminosilicate or silica molecular sieve, and an associated structural supporting component.

In the broadest aspect of the invention, the molecular sieve suitable in the packaging material may be functionally defined as any aluminosilicate or silica having a zeolite-type structure where 5 mg. of it is capable of adsorbing at least 25% of hydrogen sulfide present in a GC vial of volume 22.5 ml. at a concentration of 12,000 ppm. in 15 minutes at a temperature of 25 +/- 2 deg. C.

Alternatively, the molecular sieve may be structurally defined as an aluminosilicate or silica having a zeolite-like structure which has a three

dimensionally-connected pore-structure all of whose pores-openings have a minimum diameter of about 4 Angstroms, a Si/Al atomic ratio of 1 or greater, and has substantially all of any associated cations those of the metals of group 1A and group 2A metals.

5 The associated structural supporting component may be a polymeric or other material. It may be in the form of a film or sheet either enclosing the molecular sieve or acting as a matrix having the molecular sieve embedded in it. The supporting component must be permeable to sulfides and mercaptans. When the molecular sieve is embedded in the supporting
10 component it is present at a level of from 0.1 to 10 weight percent, preferably from 0.5 to 5 weight percent.

The packaging materials comprising the molecular sieve and the associated supporting component may be formed into a package structure itself, the packaging material forming the package walls or part of the package walls.
15 such as a layer in a laminate. Or it may be in the form of an insert such as a coupon, a cap liner and the like.

DETAILED DESCRIPTION OF THE INVENTION

In literature on molecular sieves the words or phrases 'channel', 'pore', 'pore-structure', 'aperture', 'pore-opening size' and 'pore-size' tend to be
20 used loosely, some with considerable interchangeability. In this disclosure we describe the materials as having 'channels' which form an overall 'pore-structure'. The openings through which molecules can penetrate we will refer to as 'pore-openings' having a 'pore-opening size'. They will be described as 'powders' consisting of 'particles'.

25 Molecular sieves can reasonably be said to either absorb or adsorb gases to describe the same gas take up. In this disclosure the words adsorb and adsorption are preferred.

While there are a large number of materials which can react with sulfides, mercaptans and other sulfur compounds, as well as a large number of
30 adsorbents of various types which can absorb various gases, surprisingly, in an investigation of many materials and adsorbents tried as sulfide odor removers, only a few defined molecular sieves have been found which are suitable to form part of the packaging materials of this invention. Even among molecular
35 sieves, many are quite unsuitable, and only those with very specific characteristics are useful.

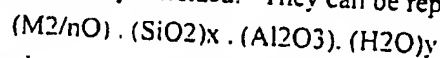
First and foremost, the molecular sieves of this invention are aluminosilicates or silicas having a zeolite-type structure. Many other metallosilicates are known, but the aluminosilicates and silicas have the major advantage of being, in general, non-toxic, and suitable for use where contact with food can or does occur. Furthermore, when blended with many polymer films, the film remains transparent, which is a major advantage in packaging films. This is because, the particle size is sufficiently small - typically, from about 0.1 to 10 microns. The lack of toxicity contrasts with packaging materials described in U.S. Patent 5,284,892, where the structure of the packaging material must prevent contact of a polyalkylene imine with food, due to its lack of acceptability in this regard. In the present invention, while it is preferred not to have small amounts of the adsorbent mixing with the food, with these materials there is no toxicity problem should they do so. Indeed, many such materials are actually used in food. Thus certain zeolites are used as moisture adsorbers admixed with table salt to allow it to pour freely.

The suitable molecular sieves of this invention are aluminosilicates or silicas having a zeolite-type structure, which are capable of absorbing a specific amount of hydrogen sulfide from an atmosphere containing it. Specifically, they are those where 5 mg. of it is capable of adsorbing at least 25% of hydrogen sulfide present in a GC vial of volume 22.5 ml. at a concentration of 12,000 ppm. in 15 minutes at a temperature of 25 +/- 2 deg. C. It is within the skill of the artisan to determine readily whether any particular molecular sieve is suitable by this simple test.

Alternatively, the suitable molecular sieve may be defined in structural terms. The critical structural parameters suitable for selecting the aluminosilicate and silica molecular sieves for the purposes of this invention are known for many such materials, but those parameters have not been determined for many others. Thus it may be immediately possible to say a given aluminosilicate or silica is suitable for use in the present invention if those parameters are known. In cases where those parameters are not known, the functional testing described above will determine suitability.

In order to describe these critical structural parameters which determined suitability, it is first necessary to describe the structure of the particular molecular sieves which are useful in the invention.

The molecular sieves which form part of the packaging materials of this invention are complex aluminosilicates or silicates characterized by a three-dimensional framework structure, enclosing cavities and channels. (These are often also referred to loosely as pores; the ends of the channels are usually described as pore-openings, or apertures, and the system of channels is said to have a pore structure). The channels may be occupied by ions and water molecules, all of which can move with significant freedom within the zeolite matrix. The water molecules can be removed or replaced without destroying the essential geometry. They may be naturally occurring, modified naturally occurring materials, or synthesized. They can be represented by the formula:



though the aluminum may approach zero, in which case the material will be referred to as a crystalline silica (not silicate). M is a cation of valence n, where the cation may be metallic or a proton. In those naturally occurring molecular sieves commonly known as 'zeolites', x is generally from about 2 to 8. And y may be 1 or more. There also exist synthesized, so-called high silica zeolites, and more generally aluminosilicate molecular sieves with extremely high levels of silica to alumina, approaching infinity; that is to say only trace amounts of aluminum are present. In that case, the material is, as noted above, essentially a silica. The cation is associated, primarily, only with the alumina, so that the less aluminum present the fewer cations. For the purposes of this invention, the term molecular sieve will be used, because the term zeolite is often defined as a material having a specific amount of aluminum, and at very low aluminum levels, approaching trace amounts, most common usage of the term zeolite would exclude such materials.

In naturally occurring zeolites, M is Na, Ca, K, Mg and Ba in varying proportions. The cations can be changed by ion exchange, and other ions such as Cs or Sr can be introduced. The structure consists of corner-linked tetrahedra with Al or Si atoms at the centers and oxygen at the corners. The tetrahedra are combined into repeat structures of 4 to 12 membered rings resulting in a framework with regular channels and 'cages' at channel intersections. The channels may be one dimensional (i.e., tubular) with the tubes not connected, two dimensional where channels branch to other channels, but each series of two dimensional channels is not connected, or a three dimensionally-connected system of channels. Here molecules which are

sufficiently small can travel freely throughout the channel system or pore-structure. In some cases there are two separate interpenetrating (but not connected) three dimensionally-connected systems. The channels and cages impart a pore-structure with differing dimensions which are critical to
5 movements of molecules into and within the channels and cages, the pore-openings controlling access to the interior of the zeolite. The pore-opening dimensions are determined not only by the tetrahedra forming the pore-opening, but by the ions in or near the pore-opening. In many cases the pore-opening is aspherical or generally asymmetrical, and can have a smallest and a largest
10 diameter. The smallest diameter will clearly control access of molecules. Overall, the variations within the zeolite structure create a wide variation in utility for various zeolites. The same situation is true for molecular sieves even when the amount of aluminum approaches just trace amounts.

The molecular sieves suitable for this invention, have a Si/Al
15 atomic ratio which can range from 1 to infinity. By infinity is meant none or only trace amounts of aluminum are present. The ratio may be determined by standard wet chemical methods or by atomic adsorption spectroscopy. In describing this type of material, it is common to use the ratio $\text{SiO}_2/\text{Al}_2\text{O}_3$ rather than Si/Al, since this characterizes the chemical units which make up the
20 structure. However in this disclosure, the Si/Al atomic ratio is used. It is of course readily possible to calculate one from the other.

The molecular sieves of this invention also have a three dimensional pore-structure, and a minimum pore-opening size of about 4 Angstrom. Pore-opening size may be somewhat temperature dependent, and
25 since adsorption in end use has to take place at ambient temperatures, the pore-opening size refers to the size which will be present at ambient temperatures, that is to say between about 20 and 30 degrees C. Methods of determining and calculating pore-opening size are well known in the art, and many molecular sieves have established pore-opening sizes.

30 Useful references generally relating to zeolite-type structure and characterization include the following: Meier et al., 'Atlas of Zeolite Structure Types' (International Zeolite Assn. 1978); Mumpton, 'Natural Zeolites' in Reviews in Minerology 14:1 (1977), and Smith, 'Origin and Structure of Zeolites' in Zeolite Chemistry and Catalysis, ACS Monograph 171,
35 American Chemical Society, 1976.

If the pore-opening is asymmetrical, the 4 Angstrom minimum refers to the smallest diameter. In addition, the molecular sieve must be basic, in that substantially all (i.e. more than about 90%) the associated cations can not be hydrogen ions, but must be a metal ion of group 1A or 2A, preferably Na, Ca, Ba, Mg and K, preferably Na and Ca. Ion-exchange with other metals such as zinc (group 2B) are also possible. As the amount of aluminum approaches zero, the number of associated cations approaches zero, and the material will become increasingly neutral. Molecular sieves falling within these limitations are suitable.

The molecular sieve's pore-opening size will be large enough to let through the smallest odor causing sulfur compound molecules. The pore-opening size of molecular sieves is a fairly precisely definable quantity. The diameter of molecular species which are to be adsorbed however is more difficult to define. Various types of measurement have been used to determine diameter, and various methods of calculating different aspects of the size of the molecule have been used. When a molecule is unsymmetrical, as in longer chain mercaptans, it becomes difficult to know what dimension of the molecule is important in relation to its ability to enter a molecular sieve having defined minimum pore-opening size. There is a crude correlation between pore-opening size and molecular size. Hydrogen sulfide has a Lennard-Jones 'kinetic' diameter of 3.6 Angstroms and a Pauling length by width of 4.6×3.9 Angstroms. These numbers are in line with what has been found to be the minimum pore-opening size allowable for adsorption of hydrogen sulfide, namely a pore-opening size of about 4 Angstroms.

Hydrogen sulfide is the principle odor causing compound from foods which generate sulfur compounds. Other sulfur containing odor causing compounds of importance are slightly larger in size, but some have a much lower odor threshold than hydrogen sulfide. If we use a value of 4.23 Angstroms for the diameter of hydrogen sulfide, one calculation/measurement/estimate of the size of various odor-causing materials made by the inventors suggested the following sizes (in Angstroms): methylmercaptan 4.80; ethylmercaptan 5.22; dimethyltrisulfide 5.43; dimethylsulfide 5.63; and methional [3-(methylthio)propionaldehyde] 6.17. The odor threshold however is in quite a different order. Thus values quoted in 'Food Chemistry' by H.D. Belitz et al. 1987, gives the following values in

micrograms per liter: hydrogen sulfide 5; methymercaptan 0.02; ethylmercaptan 0.008; dimethyltrisulfide 0.01; dimethylsulfide 1.0; and methional 0.2. While, as will be described below, sieves with a minimum pore-opening of about 4.0 Angstroms adsorb hydrogen sulfide provided the other structural parameters are met, for adsorption of other odor-causing sulfur compounds, it is preferred to have a pore-opening size of greater than about 5.5 Angstroms.

There is no particular particle size limitation for the molecular sieve. However, for adequate surface area for rapid adsorption, it is preferred that the particles be from 0.1 to 5 microns. When molecular sieve particles are incorporated into a film layer of a package (i.e., as a filler), the particles should preferably have an average diameter less than half the film thickness, and a size distribution such that no particles have a diameter as thick as the film, and preferably less than three quarters of the thickness of the film. This represents no real problem, since a 1 mil. (25 micron) film could have particles, by this criterion, of 12 microns in diameter which is higher than typical sieve particles. The adsorbent layer may be from about 1 to 5 mils.

The particulate molecular sieve adsorbent obviously is not mixed directly with the food, and is kept separate from it. The adsorbent is part of a packaging material, the packaging material comprising the adsorbent and an associated structural supporting component. By this, is simply meant another material which contains (i.e. as a filler), encloses, or otherwise supports the particulate adsorbent so that the individual adsorbent particles can not readily separate from the packaging structure and mix with the food. Thus, the adsorbent may be melt blended, just as with any particulate filler, into a polymer which is then made into a film which forms the walls or part of the walls of a packaging container. Thus the package may be a film pouch. The pouch of course may be further enclosed in some other package such as a carton.

The packaging material of the invention when it is in the form of an adsorbent-filled film may be one layer of a multilayered film or sheet, with other co-extruded or laminated and adhered layers acting as water or oxygen barriers etc. If there are other layers, the adsorbent-filled layer may be the layer which contacts the food, unlike the situation where the adsorbent is not suitable for food contact such as polyethylene imine aldehyde scavengers. It may however be another layer, provided any inner layers (those between the adsorbent layer and the food) are permeable to the odor-causing sulfur

compounds, so that they may pass through it to contact the adsorbent-filled layer. For instance the inner layer could not be a metallic film layer or highly crystalline polymer or poly(vinylidene) chloride which is a useful barrier. If there is an inner layer, it should not be more than 3 mils thick.

5 Alternatively, the package material may not be part of the package itself, but may be a free insert, such as a coupon. Or it may be an insert which does form part of the package, such as a cap liner. When the packaging material is a free insert, a possibility within the bounds of the present invention is that of the particulate adsorbent enclosed, i.e. encapsulated by a polymer film or other material, but not blended (i.e., as a filler) within it. Other materials
10 which could be used to encapsulate molecular sieve particles are non-woven materials. These would act as ideal encapsulating materials since they are highly porous, provided the pores of the non-woven are not of a size which would let through the molecular sieve particles.

15 Polymers suitable as the associated structural supporting component are olefinic homopolymers such as polypropylene, polyethylenes such as LDPE, LLDPE and ULDPE, copolymers of ethylene and vinyl esters such as vinyl acetate, and copolymers of ethylene and unsaturated acid or esters of those acids such as acrylic or methacrylic acid, or 1-8 carbon alkyl acrylates and methacrylates, or mixtures of these comonomers. Also included are
20 ionomers of ethylene/acrylic acid or methacrylic acid copolymers and terpolymers. Ionomers are the well known metal ion partially neutralized ethylene/(meth)acrylic acid copolymers, described in U.S. 3,264,272 (Rees) which is hereby incorporated by reference. The polymers preferred as the
25 structural supporting component are polyolefines and polyolefin copolymers (i.e. with other alpha olefins).

In multilayer films used to form packages, one layer of which is the packaging material of the present invention, other layers may be an inner seal layer and outer barrier layers or merely outer bulk layers. Such multilayer
30 films are well known in the packaging industry and there are endless variations of such films. The adsorbent layer is preferably separated from the food to be packaged only by one layer, that layer being a seal layer which contacts the food. Seal layers may be any of the polymers listed above as the structural supporting component, but preferably ionomers or ethylene vinyl acetate which
35 are ideal adhesives. Other layers may be an oxygen barrier layer such as

ethylene/vinyl alcohol copolymers, or a moisture barrier layer which may also be the bulk layer. Suitable polymers for this include polypropylene, polyethylene and polyethylene copolymers. The layers may have various adhesives tying them together, such as 'BYNEL' which is the name for certain modified polyolefins manufactured by E.I. du Pont de Nemours and company.

EXAMPLES

The invention is illustrated by the following examples.

Test Methods

10 Different potential reactants or adsorbents were evaluated as hydrogen sulfide removers using GC. Either the reactant/adsorbent itself, or the reactant/ adsorbent melt blended with a polymer at a level of 2 parts per 100 parts polymer were placed in a GC vial. Weight of the sample to be tested was 5 mg. when straight powdered adsorbent was tested, and about 8 square inches
15 (4 X 2 inches) of film when film was tested.

Then enough hydrogen sulfide to provide 12000 ppm in the (air containing) vial (based on the volume of the vial which was about 21.5 ml.) or about 250 uL of pure hydrogen sulfide was injected into the vial. After fifteen minutes, the concentration of hydrogen sulfide remaining in the vial was
20 measured. The concentration change was measured by taking out 250 uL of the mixed air/hydrogen sulfide from the vial, and measuring the hydrogen sulfide concentration using standard GC methods. The adsorption was calibrated against a similarly filled vial without adsorbent, to allow for adsorption by the vial septum and vial surface.

25 In the following examples, in Series 1, samples of the potential adsorbent were melt blended with 'ALATHON' 1540, a LDPE, manufactured by the Du Pont Co. at a 2 weight percent level, placed in a vial and evaluated as indicated above. A list of the potential adsorbents/reactants tried is shown in Table 1. Results are shown in Table 2. In Series 2, potential
30 reactant/adsorbent was placed directly in the vial and evaluated as indicated above. Results are shown in Table 3. Some potential reactant/adsorbents were tested both in series 1 and 2. In other cases, if a reactant/adsorbent showed no potential in one or the other of the two series, (whichever was tried first with the material) it was not repeated in the other series.

Zeolite M5 (also known as Mordenite) has the required pore-opening sized, but has a one dimensional pore structure. That is to say it has unconnected single channels. Chabazite has a three dimensionally-connected pore structure, but the pore-openings in one of the three dimensions is less than 4 Angstroms. DAY-55 has the necessary pore size, the channels are connected into a three dimensionally-connected pore-structure, but it is acidic, having protons replacing metal ions. Such acidic materials also have the disadvantage of possible reaction with some of the polymeric packaging components of the packaging material. Zeolites 4A and 5A adsorbed some hydrogen sulfide alone, but not when incorporated into film. The minimum pore-opening is defined as about 4 Angstroms. In any particular sieve the actual minimum may be from 3.7 to 4.3 Angstroms. It is surmised that the effective diameter for these zeolites under test conditions may be smaller than that quoted, at least in one direction. The zeolites with much larger pore-opening size, Y52 and 13X are clearly better and preferred.

Clays have a laminar structure, not a Zeolite-type three dimensional structure. The various clays tried generally do not perform well in films. They do typically contain a complex of Al_2O_3 and/or SiO_2 and/or MgO , but the structure is inadequate. Bentonite does perform well when tested neat, but not in films. This may be related to interplanar spacing, between 'molecules', that spacing disappearing when blended with films. By contrast, the 3 dimensional structure of the zeolite-like structures remain essentially intact when blended with polymers.

TABLE I

ABSORBENTS/REACTANTS EXAMINED

SUBSTANCE		TYPE	DIMENSIONALITY	PORE Diameter	Si/Al
ION					
5	Activated alumina		NA	NA	NA
	ZnO/TiO ₂ (1)		NA	NA	NA
	ZnO?Zn phosphate (1)		NA	NA	NA
	ZnO/Al salt/Ni salt (1)		NA	NA	NA
	Mg Aluminate		NA	NA	NA
10	Activated Carbon		NA	NA	NA
	VEEGUM A132A		NA	NA	NA
	VAN GEL B		NA	NA	NA
	Bentonite NF	A Clay	NA	NA	NA
	Bentonite MPS	A Clay	NA	NA	NA
15	Magnebrite HS	A Clay	NA	NA	NA
	Hectabrite DF	A Clay	NA	NA	NA
	Laponite RD	A Clay	NA	NA	NA
20	Chabazite (CHA type Zeolite)		3	3.8 x 3.8	2 Ca
	ABSCENTS 3000 (MFI Sieve)		3	5.3 x 5.6	>100 ?
	Zeolite 5A (LTA Zeolite)		3	4.1 x 4.1	1 Ca
	Zeolite 4A (LTA Zeolite)		3	4.1 x 4.1	1 Na
	Sieve S115	(Same as Abscents 3000? Different particle size?)			
25	Zeolite DAY-55		3	7.5?	55 H
	Zeolite M5 (MOR Zeolite)		3 / 1	6.5x 7/2.6x5.7	5 Na
	M. Sieve ETS-4	A titanosilicate	?	<4.0	NA ?
	M.Sieve ETS-10	A titanosilicate	?	8-10	NA ?
30	Zeolite Y52	FAU Zeolite	3	7.4 x 7.4	2.5 Na
	Zeolite 13X	FAU Zeolite	3	7.4 x 7.4	1.25 Na

NA NOT APPLICABLE: ? = Not Known.

ZnO/TiO₂ is Zn₂TiO₄-Zn₂Ti₃O₈ from Alfa Inorganics.

35 (1) ZnO/Zn phosphate calcinated mix as prepared in U.S. Patent No. 5,219,542

ZnO/Al salt/Ni salt calcinated mix as prepared in U.S. Patent No. 5,174,919

Magnesium aluminate calcinated as prepared in U.S. Patent No. 5,229,091

Zeolites 4A, 13X and Y52 from Alfa Inorganics.

Zeolite 5A and M5 (Mordenite) from Linde

40 Abscents 3000 and S115 from UOP Corp.

Molecular Sieves ETS-4 and ETS-10 from Englehard Corp.

Chabazite from Minerals Research, mined in Christmas AZ.

All Clays from American Colloid Company.

Activated Carbon, grade S-5 from Darco.

Zeolite DAY-55 from DEGUSSA. Corp.

VEEGUM A132A is purified smectite clay (hydrated magnesium aluminum silicate) from R.T.Vanderbilt Co. ; VAN GEL B is smectite clay from

5 Vanderbilt.Co. They are used in the oil industry to adsorb/react with sulfur compounds.

Zeolite three letter designations are those of the International Zeolite Association.

TABLE 2
ABSORPTION OF H₂S BY POLYMER BLENDS

	<u>SUBSTANCE</u>	<u>%H₂S REMOVED</u>
5	Act. alumina	0
	ZnTi oxide	0
	VEEGUM A132A	0
	VAN GEL B	0
	Mg Aluminate	0
10	ZnO-Al-Ni	29
	ZnO-Zn phosphate	0
	ABSCENTS 3000	23
	Zeolite 5A	0
	Zeolite 4A	0
15	Chabazite	0
	Act. carbon	7
	Zeolite S115	100
	Bentonite (MPS)	100
	Magnebrite HS	0
20	Hectabrite DF	3
	Laponite RD	6

All film samples contained 2 weight percent of the ingredient.

- 25 Films were extruded on a 28MM. co-rotating twin-screw Extruder, having an L/D of 28/1 and 2 mixing sections. A 10 inch film die was used to produce 1 mil film, at a melt temperature of 215-225 deg.C.

TABLE 3
ABSORPTION BY PURE ABSORBANT/REACTANT

SUBSTANCE		%H ₂ S REMOVED
5	ZnO-Zn phosphate (5.7mg)	9
	ZnO-Al-Ni (6.4mg)	100
	Zeolite 5A	13
	Zeolite 4A	17
	Zeolite 13X	100
10	Abscents 3000	54
	Zeolite Y52	86
	Zeolite M5	2
	Chabazite	4
	Zeolite DAY-55	0
15	M.Sieve S-115	40
	M.Sieve ETS-4	0
	M.Sieve ETS-10	19
	Magnebrite HS	5
	Bentonite MPS	4
20	Bentonite NF	4
	Hectabrite DF	17
	Laponite RD	6

25 Powder sample weight was 5 mg.

CLAIMS

1. A packaging material for adsorption of sulfur compound odors from foods, comprising:

5 a) a particulate aluminosilicate or silica molecular sieve having a three dimensionally-connected pore structure, having pore-openings of minimum diameter 4 Angstroms, a Si/Al atomic ratio of 1 or greater, and substantially all of any associated cations those of metals of group 1A or 2A of the Periodic Table, and

10 b) an associated structural supporting component, the associated structural supporting component ensuring that the molecular sieve is kept separate from the food to be packaged.

2. The packaging material of claim 1 wherein the associated cations are Na or Ca.

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3. The packaging material of claim 1 wherein the minimum diameter of the pore-openings is 5.5 Angstroms.

4. The packaging material of claim 1 wherein the associated structural supporting component is a matrix polymer film in which the molecular sieve is blended at a concentration of from 0.1 to 10 weight percent, and further wherein the matrix polymer film blend forms at least one layer of a film package.

25 5. The packaging material of claim 1, wherein the associated structural supporting component is a polymer film which encompasses or encapsulates powdered adsorbent, and further wherein the resulting packaging material is in a form suitable for a packaging insert.

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6. A packaging material for adsorption of sulfur compound odors from foods, comprising:

5 (a) a particulate aluminosilicate or silica molecular sieve having an adsorption capacity such that 5 mg. placed in a GC vial is capable of adsorbing at least 25% of hydrogen sulfide present in a GC vial of volume 22.5 cc., at a concentration of 12,000 ppm in the vial, in 15 minutes at 25 +/- 2 deg.C. and

(b) an associated structural supporting component, the associated structural supporting component ensuring that the molecular sieve is kept separate from the food to be packaged.

10

7. A packaging material for adsorption of sulfur compound odors from foods, comprising:

(a) a particulate aluminosilicate selected from the group comprising zeolite Y52 and zeolite 13X, and

15 (b) an associated structural supporting component, the associated structural supporting component ensuring that the molecular sieve is kept separate from the food to be packaged.

INTERNATIONAL SEARCH REPORT

International Application No.
PCT/US 96/00258

A. CLASSIFICATION OF SUBJECT MATTER
IPC 6 B01J20/18

According to International Patent Classification (IPC) or to both national classification and IPC:

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)
IPC 6 B01J

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
Y	EP,A,0 297 543 (UNION CARBIDE) 4 January 1989 see page 3, line 15 - page 5, line 58 ---	1-5
Y	DATABASE WPI Week 9420 Derwent Publications Ltd., London, GB; AN 94-164063 XP002003000 & JP,A,06 107 847 (FUJI KAGAKU KOGYO) , 19 April 1994 see abstract --- -/--	1-5

☒ Further documents are listed in the continuation of this C.

☒ Patent family members are listed in annex.

* Special categories of cited documents:

- *A* document defining the general state of the art which is not considered to be of particular relevance
- *E* earlier document but published on or after the international filing date
- *L* document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)
- *O* document referring to an oral disclosure, use, exhibition or other means
- *P* document published prior to the international filing date but later than the priority date claimed

- *T* later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention
- *X* document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone
- *Y* document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art.
- *Z* document member of the same patent family

Date of the actual completion of the international search

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INTERNATIONAL SEARCH REPORT

Inter. Application No.
PCT/US 96/00258

(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT

Category	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
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A	<p> DATABASE WPI Week 9043 Derwent Publications Ltd., London, GB; AN 90-324979 XP002003002 & JP,A,02 233 138 (TOSOH CORP.) , 14 September 1990 see abstract --- </p>	1
A	<p> EP,A,0 159 056 (SHELL) 23 October 1985 see page 9-10; claims 1-18 ----- </p>	1-3

Form PCT 15A 210 (continuation of second sheet) (July 1992)

INTERNATIONAL SEARCH REPORT

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PCT/US 96/00258

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